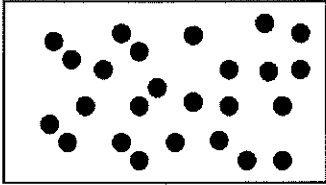
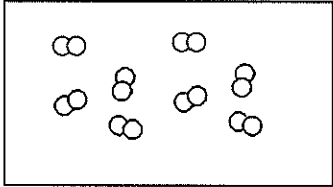
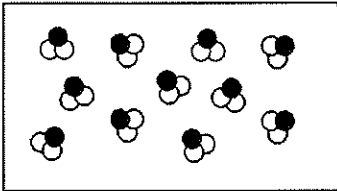
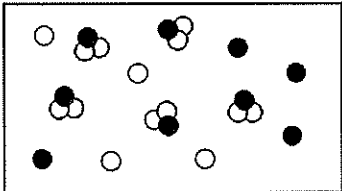


Elements Compounds and Mixtures Study Sheet

<p>Elements</p>	<ul style="list-style-type: none"> - Are a pure substance. - Cannot be broken down into simpler substances - Are found on the periodic table. - Made up of all the same atoms. <div style="display: flex; justify-content: space-around; align-items: center;">   </div>
<p>Compounds</p>	<ul style="list-style-type: none"> - Pure substance - Created in a chemical reaction. - A new substance with new properties is formed. - Connected chemically using chemical bonds. - Always has a formula (CO₂, H₂O, NaCl) <div style="text-align: center;">  </div>
<p>Mixtures</p>	<ul style="list-style-type: none"> - Not a pure substance. - 2 or more different pure substances that are physically combined. - Can be separated by physical means. <div style="text-align: center;">  </div>
<p>Homogeneous Mixture</p>	<ul style="list-style-type: none"> - A mixture that is completely mixed and looks the same throughout. <p>Examples: Milk, Orange Juice, Coffee, Tea</p>
<p>Heterogeneous Mixture</p>	<ul style="list-style-type: none"> - A mixture that is not completely mixed. You can see the different parts. <p>Examples: Chex Mix, Sandstone, Cereal</p>
<p>How to find the number of atoms in a chemical formula.</p>	<ul style="list-style-type: none"> - If the Element has a small number to the right of it that means it has that many of that atom. If there is no number next to the Element there is only one atom of that element. <p>Example: H₂SO₄ = 2 hydrogen atoms, 1 sulfur atom, 4 oxygen atoms =7</p> <p>Example: NaCl= 1 Sodium, 1 Chlorine</p>